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An acid is a substance that contains hydrogen and ionizes to produce hydrogen ions in aqueous solution; A base is a substance that contains a hydroxide group and

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dissociates to produce a hydroxide ion in aqueous solution. Bronstred-Lowry model. A model of acids and bases in which an acid is a hydrogen-ion donor and a base is a hydrogen-ion acceptor.

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Chapter 19 - Acids & Bases Chapter 19 covers properties of acids and bases, the different definitions, $[H^+]$ and $[OH^-]$ and pH and pOH, K_a and K_b and their use when dealing with weak acids and bases and buffers.

Chapter 19 - Acids &
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Bases - Mrs. Gingras' Chemistry Page

He said that acids are hydrogen containing compounds that ionize to yield hydrogen ions (H^+) in aqueous solution. Also said that bases are compounds that ionize to yield hydroxide ions (OH^-) in aqueous solution

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Answer. The
Arrhenius definition
states that acids are
compounds that

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contain hydrogen, and when placed in aqueous solution, ionize to form hydrogen (H^+ ions). Bases are compounds that ionize to yield hydroxide ions (OH^-) in aqueous solutions.

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Obtain, evaluate, and communicate information about the properties that describe solutions and the nature of acids and

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bases. f. Use mathematics and computational thinking to compare, contrast, and evaluate the nature of acids and bases in terms of percent dissociation, hydronium ion concentration, and pH.

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SECTION 19.1

ACID-BASE

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THEORIES(pages 587-593) This section compares and contrasts acids and bases as defined by the theories of Arrhenius, Brønsted-Lowry, and Lewis. It also identifies conjugate acid-base pairs in acid-base reactions.

**Chapter 19: Acids
and Bases
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Section 19.1 - Acid-Base Theories Acids have a sour taste, change the color of an indicator, can be strong or weak electrolytes in aqueous solution, and react with metals.

Chapter 19 - Acids, Bases, and Salts

of acid or base are added. A buffer system should contain considerably more acid than base. Specific

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buffer systems should be chosen based on the pH that must be maintained. 114 Study Guide for Content Mastery Chemistry: Matter and Change Chapter 19

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Lesson Worksheets

Acids have a sour, tart taste, react with metals such as Zn and Mg, react with compounds containing Hydroxide ions (good proton acceptors to form water and a salt).

Bases have a bitter taste, feel slippery, reacts with an acid to form water and a salt, and can form strong or weak electrolytes in aqueous solution. Work

Step by Step

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Arrhenius said that acids are hydrogen-containing compounds that ionize to yield hydrogen ions (H^+) in aqueous solution. He also said that bases are compounds that ionize to yield hydroxide ions (OH^-) in aqueous solution.

Chapter 19 Acids,

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634 Chapter 18 • Acids
and Bases Section

118.18.1 Figure 18.1
Rhododendrons

flourish in rich, moist
soil that is moderately
acidic, whereas semper-
vivum, commonly
called hen and chicks,
grows best in drier,
slightly basic soil.

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the behavior of acids
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